

# Ionic Bonding – Lewis Dot Structures

**Ionic Bond** – Bond between a metal and a non-metal.

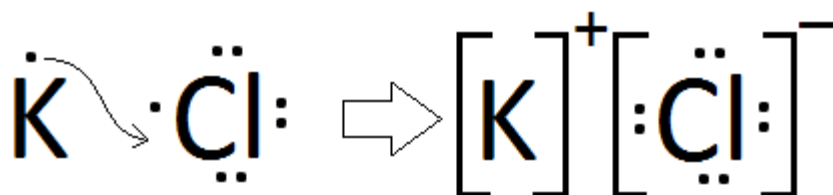
Electrons are **transferred**, not shared.



**Draw the correct Lewis Dot Structure for each of the compounds created in each question**

**Example:** K + Cl

Lewis Dot Structure:



1) Li + Cl

Lewis Dot Structure:

2) Mg + O

Lewis Dot Structure:

3)  $\text{Ca} + \text{Br}$

Lewis Dot Structure:

4)  $\text{K} + \text{O}$

Lewis Dot Structure:

5)  $\text{Na} + \text{N}$

Lewis Dot Structure:

6)  $\text{Mg} + \text{F}$

Lewis Dot Structure:

7)  $\text{Li} + \text{S}$

Lewis Dot Structure:

8)  $\text{Mg} + \text{S}$

Lewis Dot Structure:

9)  $\text{Cs} + \text{P}$

Lewis Dot Structure:

10)  $\text{Na} + \text{Cl}$

Lewis Dot Structure: